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By Leendert van der Plas (Wageningen)<sup>1</sup>) and Theodor Hügi (Bern)<sup>2</sup>)

With 2 figures in the text

#### Abstract

Chemical, optical, X-ray, DTA and TGA data are given for a Ferrian sodiumamphibole, separated from a metamorphic ophiolite in the northern Adula region (near Vals, Kt. Graubünden, Switzerland). Information regarding the following trace elements is added; Ba, Co, Cr, Cu, Mo, Ni, Pb, Sn, Sr, V, Zn. The chemical analysis is as follows: SiO<sub>2</sub> 49.22, Al<sub>2</sub>O<sub>3</sub> 7.78, Fe<sub>2</sub>O<sub>3</sub> 8.28, FeO 10.49, MnO 0.23, MgO 10.52, CaO 7.45, Na<sub>2</sub>O 3.54, K<sub>2</sub>O 0.26, H<sub>2</sub>O<sup>+</sup> 1.80, H<sub>2</sub>O<sup>-</sup> 0.32, TiO<sub>2</sub> 0.23 = 100.12. The optical properties. comparable to those of ferrohastingsite, are X 1.66—1.67, Y n. d., Z 1.675—1.685,  $2V_x$  0°—48° (axial plane normal to 010, Z = b),  $2V_x$  0°—20° (axial plane parallel to 010, Y = b). Chemical and optical data of some riebeckites, ferrohastingsites, femaghastingsites and magnesiohastingsites from other localities are tabulated.

#### Introduction

While mapping a part of the northern Adula region, an outcrop of bluish black rocks was encountered SW of the village Vals, Kanton Graubünden, Switzerland. For a description of the geological setting of these rocks, of their constituent minerals and of their probable genesis, the reader is referred to VAN DER PLAS (1959). A summary of these data is given below.

Samples LP 88, LP 88a and LP 452. W of the Hohbühl, coord. 732,6—160,4, alt. 2400 m. A metamorphic ophiolite of presumably Mesozoic age.

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#### Macroscopical description:

A massive fine-grained bluish-black rock, showing epidote patches, magnetite octahedrons and fine amphibole crystals, and cut by an albite vein containing *black amphibole* needles of about 4 cm length and not exceeding 1 mm in diameter.

#### Microscopical description:

The host rock shows patches with and epidote fabric and patches of parallel amphibole prisms with magnetite octahedrons. The size of these patches is about 2 to 3 cm and their shape irregular. The epidote patches consist chiefly of granular or fanlike yellowish epidote with an occasional glaucophane prism (0.2 mm). The patches built up of parallel prisms of a weakly coloured amphibole with magnetite hold several crystals of *amphibole* with *dark greenish blue cores* and light coloured rims. The transitional fabric between these extremes consists of the latter type of amphibole, with light coloured rims set in a matrix of albite and epidote together with separate crystals of a lighter coloured amphibole. The *deep coloured amphibole* in this transitional fabric occasionally shows cores of glaucophane, whereas the glaucophane in the epidote parts frequently occurs in direct contact with epidote.

It became apparent that dark green amphibole is of rather uncommon composition; the same mineral was found to be beautifully developed in an albite vein showing slightly fibrous rims and a less pronounced zonal habit. In the paper mentioned above, this amphibole was called "ferrohastingsite", on the assumption that similarity of optical properties might indicate a similarity of chemical composition. As ferrohastingsite has not yet been described from Switzerland it seemed worth while to carry out a mineralogical investigation. The results of these investigations are reported here.

# **Optical Properties**

Within the zonal crystals the optical properties vary. The range of this variation is as follows:

nX	1.66 - 1.67.
nY	n. d.
nZ	1.675 - 1.685.
nZ—nX	0.010-0.016.
$2V_x$ 0°-48°	axial plane normal to 010, $Z = b$ .
$2V_{\mathbf{x}}  0^{\circ} \underline{} 20^{\circ}$	axial plane parallel to 010, $Y = b$
c/Y 20°-30°	
c/Z 12°-20°	

The colour parallel to the c-axis is a pronounced blue. The colour parallel to the b-axis is olive green.

The colour approximately parallel to the a-axis is yellowish.

Measurements were all carried out in Na light. Values for other wavelengths may differ appreciably from the data listed. Absorption is less strong than with riebeckites. Extinction dispersion in daylight is large and dispersion of the optical axes extreme. Differences between the measurements made in Li light and those made in Na light in the same crystal illustrate this clearly, viz  $2V_{xLi}$  41°,  $2V_{xNa}$  48°. There are strong indications that crossed dispersion is present, but strong absorption makes observations difficult.

BILLINGS (1928) was the first to compile optical and chemical data of hastingsites. These optical data are listed together with the chemical composition in tables 4 and 5. The optical properties of femaghastingsite, magnesiohastingsite and alkalihastingsite conform less well to our measurements than those of ferrohastingsite. The refringence of the amphibole under discussion is, however, rather low for ferrohastingsite but its other optical properties fit well.

In view of the chemical composition of the investigated amphibole it seems appropriate to compare the optical properties with those of riebeckites. MIYASHIRO (1957) recently compiled data of chemically analyzed sodium-amphiboles and came to the conclusion that the literature shows considerable confusion and contains many contradictions concerning the optical descriptions of riebeckites as a result mainly of very strong absorption. Surveying these optical data and comparing them with our observations one is lead to the conclusion that the refringences of riebeckites are always higher, whereas those of magnesioriebeckite conform to those of the amphiboles we investigated. At the moment there are not enough optical data available concerning magnesioriebeckites to draw a conclusion from this fact. The optical properties of our amphibole differ so strongly from those of riebeckites that before the chemical analysis became available riebeckite was not even thought of as a main component of this amphibole. In conclusion it can be stated that the amphibole from Vals is most closely allied to the group of hastingsites as far as the optical properties are concerned. There is some slight indication that magnesioriebeckites have comparable optical properties.

# **X-Ray Investigations**

A sample, prepared with the aid of heavy liquids was used in X-ray investigations. Crystals that sank in a mixture of methylene iodide and acetone of S.G. 3.100 were collected, and when examined under a micro-

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	Ferrian		& Co	5.62	6.07	10.60	11.42	12.24	12.77	13.25	13.72	14.07	15.20	15.83	16.21	16.66		17.60	18.61		19.21	20.10	20.55		22.55	23.10	23.79	24.14	24.38	24.64

Table 1. X-ray powder patterns of ferrian sodium-amphibole and other amphiboles

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2.080	2.016	]	]	ļ	1.875	1.836	1.807	]	Į	1.689	1.665		1.644	1.618	1.593		1.517	1.504	]	1.459	1.440		1.386	1.349	1.321	1.311	1.295	1.273	1.235	1.197	1.168	1.086	1.050	1.029	1.018	0.987	0.981
27.74	28.69	Į		1	31.08	31.81	32.39	}		34.96	35.54		36.08	36.75	37.42	l	39.64	40.06	]	41.55	42.25		44.32	45.85	47.12	47.57	48.37	49.51	51.61	53.97	55.97	63.07	67.21	70.20	71.95	78.83	80.77
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2.058	2.019	1.998	1.957	1.909	1.862	1.823	1.805	1.776	1.698	1.683	1	1.653		1.614	1.587	1.557	1.523	1.502		l	1	1			]	1	[	1	]								
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2.06	2.02	2.02	1.961	1.928	1.866		1.803	1.738		1.685	]	1.650	1.632	1.617	1.589	1.562	1.520	1.501	1.474	1.457	1.440	1.430	1.368	1.345	1.318	]	1.295	1.192	cut off								
25.72	26.29	26.35	27.14	27.64	28.65		29.75	30.98	ļ	32.07	]	32.82	33.24	33.59	34.26	34.94	36.04	36.58	37.37	37.88	38.40	38.72	40.82	41.69	42.75	1	43.70	48.64						53 			

scope revealed that the light-coloured amphibole variety, forming rims around the more deeply-coloured cores, was present only in very small quantity. It has to be kept in mind, however, that the amphibole crystals on which our observations were made are still slightly zonal.

Part of this sample was put in a Lindeman glass tube and investigated with Co K  $\alpha$  radiation using a powder camera,  $\emptyset$  114,6 mm, with a collimator with rectangular aperture. The camera was filled with H<sub>2</sub> gas



Fig. 1. Ferrian sodium-amphibole from Vals. Co K<sub>α</sub> radiation.
a) Ø 114.6 mm b) Ø 57.3 mm

under normal atmospheric pressure and the film mounted according to the conventional Straumanis-Ievins arrangement. The diameter of the diffraction rings was measured with a low-power travelling microscope. It is our experience that d-values, thus obtained from measurements on a film, are superior to those obtained with a diffractometer unless the diffractometer is very carefully aligned and measurements are made manually on counters all of which is rather time-consuming. In table 1 the d-values are listed as well as the values of  $\theta_{\rm Co}$ . The pattern is indexed according to JAKOB and BRANDENBERGER (1931) as far as this coincides with the indexing of Perrault. As the paper of PERRAULT, announced for the Can. Min., 6, part 3, 1959, is still to appear later in 1961, his values have been taken from card 9-434 of the ASTM index. The powder patterns of two riebeckites are given for the sake of comparison.

Another part of the same sample was carefully ground in an agate mortar. The powder was stirred in a 0.005 N NaOH solution and left to

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settle for the time prescribed for a given temperature and height of the liquid column. After the grains larger than 8  $\mu$  has settled, the supernatant liquid was siphoned off and the fraction  $< 8 \mu$  was collected with the aid of a centrifuge. The fraction  $< 0.5 \mu$  was largely disposed of by washing three times with H<sub>2</sub>O in a centrifuge and decanting the still somewhat turbid liquid each time. In this way one is sure to get rid of the extremely small crystals that might otherwise cause a broadening of reflections. The sample thus prepared was used in determining the relative intensities of the reflections with the aid of a Hilger and Watts X-ray diffractometer type Y 125 with scintillation counters. This diffractometer is of the monitoring type.

To avoid as much as possible the effects of preferred orientation, three different techniques were used to prepare the sample surface; i: the surface was pressed on a ground glassplate; ii: the surface was pressed on a piece of filter-paper; iii: two tablets were produced in a small hand-press and the surfaces of these tablets ground against each other. After each run the sample was prepared anew. As tests for the best method of preparing unoriented samples of material that might give rise to preferred orientation are still in progress, comment on the characteristics of the three methods would at present be premature. It suffices to report that the results in this case were more or less comparable with respect to reproducibility. On the other hand, samples with a grain-size  $< 50 \mu$  gave highly unreliable results, i.e., a few peaks with relative intensities < 80 in samples with a grain-size  $< 8 \mu$  practically disappeared in some of the former traces. The runs made with different scanning speeds were started at  $2\frac{1}{2}^{\circ} \theta$  and cut off at  $45^{\circ} \theta$ , or vice versa. The most important peaks were scanned manually on duo-gate counters. The relative intensities of the different diffraction lines are listed in table 1.

## DTA and TGA

A DTA was made in normal atmosphere with a heating rate of  $8^{\circ}$ /min. The trace shows three important exothermic reactions occurring at  $670^{\circ}$ ,  $975^{\circ}$  and  $1090^{\circ}$ . The last one is most certainly due to a collapse of the amphibole structure followed by total recrystallization. X-ray investigation of the structure of amphiboles directly after these exothermic reactions was left for subsequent investigation.

A TGA was made in N<sub>2</sub> atmosphere with a heating rate of  $2\frac{1}{2}^{\circ}/\text{min}$ . After the maximum temperature of the apparatus, viz 1060°, was reached, the gas supply was cut and the sample was left for one hour at this maximum temperature in a normal atmosphere in order to oxidize the available FeO. Inspection of the sample after cooling, however, showed that total oxidation did not take place. By this method we had hoped to obtain a check on the amount of FeO found in the chemical analysis. The amount of water lost at  $1060^{\circ}$ C was 1.73%. It is seen in fig. 2 that the TGA trace shows a change of gradient shortly after the important exothermic or endothermic reactions, illustrated by the DTA trace, have taken place. The small secondary reactions accompanying the important exothermic and endothermic reactions are partly due to the technique applied, partly to other influences not to be discussed here.

It needs to be mentioned that the DTA trace is plotted on a scale of which unity along the y axis, is twice as large as along the temperature axis.





### **Chemical Composition**

Chemical and spectrochemical analyses were carried out on a 99% pure sample of amphibole crystals prepared in the same way as the sample for X-ray analysis by using heavy liquids. The results are listed in table 2.

The analysis (column a) was carried out by means of classical methods. Values of the same sample LP 88 a given in column b) are based on complexing agents (for Ca and Mg) and flame-photometer-methods (for Na and K) and spectrographic methods (for Mn and Ti). The heavy metals have been eliminated in order to get better titrimetric results for Ca and Mg by means of "Komplexon III". We used the HHSNNand "Eriochromschwarz T"-indicator for Ca and Mg respectively (SCHWARZENBACH 1960) and an EEL-flame-photometer for the alkalis. The differences for these values listed in column a) and b) lie between 0.6% and 4% (for K<sub>2</sub>O). MnO and TiO<sub>2</sub> have been checked by means of spectrographic methods (Hilger — automatic large quartz and glass spectrograph E 478, interrupted D.C. carbon-arc, 8 or 12 A, RSVinstrument, Dr. R. Seitner and Co., Type FES/240 GTT 5. Jarrell-Ash microphotometer comparator, Type JA 2100). Table 2a gives some additionel semiquantitative spectrographic results. The H<sub>2</sub>O<sup>+110</sup>°-value has been controlled by means of T.G.A. in a Nitrogen atmosphere (see p. 378).

Table 2. $Am_{f}$	phibole from 1	Vals, sample	. <i>LP 88 a</i> ; W	of the H	ohbühl; coord.
732,6-160,4	, alt. 2400 m	; S. G. 3.31;	analysts TH.	Hügi ar	nd E. Spycher

	a)	b)		c)	<b>d</b> )	e)
			Kation a	aequ. Kati	ons / 23 O	total positive
	Wt %		[based or	n data of c	olumn a)]	charges
SiO <sub>2</sub>	49.22		Si	8191	716	<b>28.64</b>
$Al_2O_3$	7.78		Al	1526	133	3.99
$Fe_2O_3$	8.28		$\mathrm{Fe}^{+3}$	1036	91	2.73
FeO	10.49		$\mathrm{Fe^{+2}}$	1461	128	2.56
MnO	0.23	0.24	Mn	32	3	0.06
MgO	10.52	10.58	$\mathbf{M}\mathbf{g}$	2609	$\boldsymbol{228}$	4.56
CaO	7.45	7.28	Ca	1328	116	2.32
Na <sub>2</sub> O	3.54	3.64	Na	1142	100	1.00
K2O -	0.26	0.22	K	55	5	0.05
$H_2O^+$	1.80	1.73				
$H_2O^-$	0.32					
$TiO_2$	0.23	0.21/0.22*)	$\mathbf{Ti}$	28	3	0.12
F	n. d.					
	100.12	*) det. H. S	CHWANDE	R	1523	46.03

This analyses leads to the following formula:

 $\begin{array}{l} (\mathrm{Ca_{1.16}\ Na_{1.00}\ K_{0.05}})^{[8]}_{2.21}\ (\mathrm{Al_{0.49}\ Fe^{+3}_{0.91}\ Fe^{+2}_{1.28}\ Mg_{2.28}\ Mn_{0.03}\ Ti_{0.03}})^{[6]}_{5.02}}\\ (\mathrm{OH},\ \mathrm{F})_2\ (\mathrm{Si_{7.16}\ Al_{0.84}})^{[4]}_8\ \mathrm{O}_{22}. \end{array}$ 

The ideal amphibole formula has 46.00 positive charges. In order to check the accuracy of the foregoing analysis, the charges of the different kations have been added proportionally in the last column and the total shows a difference of only 0.03.

Table 2a. Trace element contents of the ferrian sodium-amphibole form Vals(all values in ppm)

Ba	<b>27</b>	or	BaO	31
Co	6	,,	CoO	8
$\mathbf{Cr}$	22	,,	$Cr_2O_3$	29
Cu	87	,,	CuO	110
Mo	12	,,	MoO <sub>3</sub>	18
Ni	<b>20</b>	,,	NiO	<b>25</b>
Pb	4	,,	PbO	9
$\mathbf{Sn}$	9	,,	$SnO_2$	11
$\mathbf{Sr}$	47	,,	$\mathbf{SrO}$	55
V	193	,,	$V_2O_5$	345
$\mathbf{Zr}$	9	,,	$ZrO_2$	13

Table 2b. Semiquantitative Fluorescent X-ray spectrographic analysis of ferrian sodium-amphibole, from Vals, sample LP 1,5/7 60; W of the Hohbühl; coord. 732,6—160,4, alt. 2400 m; analyst R. O. MÜLLER

	Method 1	Method 2	Chemical analysis Column a) Table 2
ZnO	ca. 0.06 %	0.045 - 0.100 %	
CuO	ca. 0.015 %	0.012 - 0.029 %	
NiO	ca. 0.008 %	0.0075-0.016 %	—
$V_2O_5$	ca. 0.07 %	0.05 - 0.13 %	
$TiO_2$	ca. 0.115 %	0.097 - 0.245 %	0.23 %
MnO	ca. 0.38 %	0.23 - 0.66 %	0.23 %

According to the analyst, Dr. R. O. MÜLLER, these values are only approximations and the results are therefore not to be treated as absolute but rather as indications of the real values. For comparison the values of  $\text{TiO}_2$  and MnO were also determined in order to test the accuracy of this semiquantitative method. In both cases the figures differ considerably from those found by the first method but fall within the range given by the second method. To comment on the methods themselves is beyond the scope of this paper and it suffices here to state that the values in the first and second columns are the result of two different determinations and calculations based on essentially different theories.

From a mineralogical point of view it is interesting that Zn and V can be present in amounts as large as approximately 0.1%. As comparable data on the presence of these elements in metamorphic amphiboles is lacking no further comments will be made in this respect.

The composition of the amphibole corresponds to two possible mixtures of pure endmembers of which only the following are possible.

Titaniferous amphibole	$Ca_2(Mg, Fe^{+2})$ Ti	$\mathrm{Si_6Al_2O_{22}(OH)}$	, F)2
Riebeckite	$\mathbf{Na_2Fe_2^{+3}Fe_3^{+2}}$	Si <sub>8</sub> O <sub>22</sub> (OH	, F)2
Glaucophane	$Na_2Al_2(Mg, Fe^{+2})_3$	Si <sub>8</sub> O <sub>22</sub> (OH	, F)2
Ferrohastingsite	Na Ca <sub>2</sub> Al Fe <sub>4</sub> <sup>+2</sup>	$\mathrm{Si_6Al_2O_{22}(OH)}$	, F)2
Ferrian hastingsite	Na Ca <sub>2</sub> Fe <sup>+3</sup> (Mg, Fe <sup>+2</sup> ) <sub>4</sub>	$\mathrm{Si_6Al_2O_{22}(OH)}$	, F)2
Tschermakite	$Ca_2Al_2(Mg, Fe^{+2})_3$	Si6Al2O22(OH	, F)2
Tremolite	$\operatorname{Ca}_2(\operatorname{Mg},\operatorname{Fe}^{+2})_5$	$\operatorname{Si}_6$ $O_{22}(\operatorname{OH})$	, F)2

The method of assigning the available kations to the endmembers is shown in the following scheme.

	Titaniferous amphibole	Riebeckite and glaucophane Riebeckite	Ferrian- hastingsite Ferrian- and ferro- hastingsite	Tschermakite	Tremolite	Rest
		I or II	I or II			
Si 716	18	336	126	105	128	3
Al 84	6		42	36		· _ ,
Al 49		14	14	35		_
Fe <sup>+3</sup> 91		70 84	21 7			
Mg 228	10	100		59	00	
$\begin{array}{c c} \mathbf{Fe}^{+2} & 128 \\ \mathbf{Mn} & 3 \end{array}$	12	120	84	53	80	4
Ti 3	3	_				
Ca 116	6		42	36	32	
Na+K 84	, Ť	84		_		
Na+K 21			21			—

To take full account of the available Ti, the total amount of Ti is assigned to titaniferous amphibole. The high amount of Fe<sup>+3</sup> points to at least the presence of a ferri-amphibole and can be assigned to riebeckite in combination with the available amount of Na in 8-fold coordination, on the assumption that kations in 10-fold coordination (cf. HERITSCH et al. 1957), viz 0.21, form the hastingsite molecule, and must necessarily be sodium. We tried FUMIKO SHÎDO'S variant (1958), assuming that some of the Ca is present in form of  $Ca_2Ca_5Si_8O_{22}(OH)_2$ . This did not provide a solution, however, which proves that the kations in 10-fold coordination are all sodium. If, however, the hastingsite is in the form of a ferrian hastingsite then the fact that the Fe<sup>+3</sup> content is insufficient to match the amount of Na in 8-fold coordination, would mean that some glaucophane is present. Both possibilities are listed as II and I respectively. What is left of the Al in 6-fold coordination, viz 0.35, is grouped into the Tschermakite molecule. The remaining ions are used to form tremolite, leaving 0.03 Si and 0.04 of the divalent metal ions as a remainder.

Summarizing we find:

3% titaniferous amphibole 35% or 42% riebeckite 7% or 0% glaucophane 21% or 7% ferrian hastingsite 0% or 14% ferrohastingsite 18% tschermakite 16% tremolite

The question as to which of the two possibilities seems to be more in accordance with the optical and X-ray data will now be discussed. The optical data are definitely in favour of the highest possible amount of ferrohastingsite with the lowest possible amount of riebeckite. This leads to the following composition:

3%	titaniferous amphibole
35%	riebeckite
7%	glaucophane
7%	ferrian hastingsite
14%	ferrohastingsite
18%	tschermakite
16%	tremolite

The X-ray powder pattern shows the same characteristics as the pattern of riebeckite, i. e., the presence of four weak lines at approximately d=1.66; 1.64; 1.62 and 1.59 Å. The patterns of Ca-amphiboles and glaucophane show a different configuration of spacings as was shown by ZWAAN and VAN DER PLAS (1958). From a structural point of view the

composition has to be such as to take into account the presence of a large amount of riebeckite. A lack of X-ray powder data of chemically analyzed sodium-amphiboles makes it impossible to estimate what percentage of the riebeckite molecule suffices to account for the presence of this typical pattern.

One problem not often treated in papers on amphiboles remains. This concerns the structure of a zonal amphibole. One must imagine a more or less statistical spread of the available kations over the specific sites, showing a gradient of concentration of one over the other from core to rim, in other words a higher concentration of e. g., sodium in a certain diffuse zone in the core gradually giving way to a higher concentration of e. g., calcium in a zone near to the rim? Or has such a zonal crystal to be treated as a multiple shell comprising layers of homoaxial crystals, each layer having a discrete chemical composition and arranged in such a way that within the amphibole the chemical composition changes discontinuously from one shell to the next?

According to the last view a powder pattern involving a mixture of patterns of homogeneous amphiboles is theoretically to be expected. As present knowledge of the X-ray powder pattern of amphiboles is very limited, and as hitherto no attempt appears to have been made to compare the pattern of a zonal amphibole with that of a mixture of pure endmembers, this question must for the moment remain unanswered. Investigations along these lines will not only throw light on the structure of zonal amphiboles, but will also be highly interesting in the light of the metamorphic facies problem.

### **Concluding Remarks**

After examining the ferrohastingsite-like amphibole from Vals, the problem remains of giving it an appropriate name. In the following three tables 3, 4, 5, the chemical and optical properties of a number of amphiboles with certain comparable properties are given. Table 3 lists the chemical composition of a number of riebeckites. From this table it is clear that these amphiboles differ considerably from our amphibole with respect to the contents of  $Al_2O_3$ ,  $Fe_2O_3$  and CaO. The next table, 4, shows a number of chemically and optically analyzed ferrohastingsites. From this table it can be seen that our mineral has too little  $Al_2O_3$  and FeO. Calculations of the composition in terms of endmembers from which the first five analyses could be derived (BILLINGS analyses) lead to at least 45% of hastingsite molecules in all these cases. This is a much

.80	49.22	7.78	8.28	10.49	0.23	10.52	7.45	3.54	0.26	0.23	n. d.	1.80	0.32	n. d.		R		100.12
7	55.44	0.22	16.77	5.23	0.09	12.30	2.17	6.76	0.15	0.04	n. d.	0.55	0.05	0.18	CI 0.08	100.03	0.09	99.94
9	55.62	4.54	12.99	3.53	1.25	11.98	1.95	5.58	0.36	0.26	0.07	1.96	0.00	n. d.				100.09
ũ	51.3	2.7	14.2	18.5	0.9	n. d.	1.1	6.0	1.4	1.9	n. d.	1.9	0.1	0.3	100.3		0.1	100.2
4	48.28	0.73	14.23	16.98	0.53	3.69	2.26	4.30	0.77	3.91	n. d.	3.32	0.36	n. d.			- 100 - 100	99.36
ŝ	47.60	3.49	11.06	18.79	0.99	4.56	2.56	4.72	1.42	1.09	n. d.	2.84	0.74	n. d.				99.86
63	52.41	0.61	14.37	14.82	1.46	5.07	1.33	4.94	2.10	0.45	n. d.	2.02	0.10	0.30	99.98		0.13	99.85
1	51.94	0.20	18.64	19.39	n. d.	1.37	0.19	6.07	0.04	n. d.	n. d.	2.58	0.31	n. d.				100.73
	$SiO_2$	$Al_2O_3$	$\rm Fe_2O_3$	FeO	MnO	MgO	CaO	$Na_2O$	$K_2O$	$TiO_2$	$P_2O_5$	$H_2O^+$	H <sub>2</sub> 0-	£.	Sum	2	FI 0 ₩ 11	

Table 3. Chemical analyses and optical properties of some riebeckites

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Z-X	0.008	n. d.	0.011	0.002	0.016	0.010	n. d.	0.010-0.016
X	1.698	n. d.	1.680	1.688	1.701	1.660	n. d.	1.66 - 1.67
Υ	1.699	1.686	1.687	n. d.	1.710	n. d.	n. d.	n. d.
Ζ	1.706	n. d.	1.691	1.691	1.717	1.670	n. d.	1.675 - 1.685
X	indigo	n. d.	n. d.	n. d.	indigo	pale yellow	n. d.	c = blue
y	yellow	n. d.	n. d.	n. d.	brownish yellow	blue .	n. d.	$\mathbf{b} = \mathbf{olive} \ \mathbf{green}$
2	indigo	n. d.	n. d.	n. d.	indigo	purple	n. d.	a = yellowish
ext.	$X/c 0^{\circ}$	n. d.	n. d.	$X/c \frac{1}{2} - 1 \frac{1}{2}$	X/c 3°-4°	$\tilde{\rm Y}/{ m a}$ 28°	n. d.	$Y/c 20^{\circ}-30^{\circ}$
2 V	z = 2	n. d.	n. d.	$z = 65^{\circ}$	$\mathbf{x} = \mathbf{ca.} 80^{\circ}$	$\mathbf{x} = 43^{\circ}$	n. d.	$Z/c  12^{\circ}-20^{\circ}$
								$\mathbf{x} = 0^{\circ} - 48^{\circ}$
disp.	ę.	n. d.	n. d.	r > v	n. d.	$\mathbf{r}\!\ll\!\mathbf{v}$	n. d.	
p =	$\mathbf{b} = \mathbf{Z}$	n. d.	n. d.	$\mathbf{b} = \mathbf{Z}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Z}$	n. d.	$\mathbf{b} = \mathbf{Z}$ or $\mathbf{Y}$
S.G.	n. d.	n. d.	n. d.	n. d.	3.32	n. d.	n. d.	3.31
	n. d.	n. d.	n. d.	$\mathbf{Z} > \mathbf{X} > \mathbf{Y}$	n. d.	n. d.	n. d.	c > b > a
					а		,	
No. 1.	Crocidolite, Kli	iphuis, analyst	H. E. VASSAR	(PEACOCK 1928	8, p. 256, No. 1).			
No. 2.	Riebeckite fro MIYASHIRO, T.,	m aegirine-riek 1956, p. 48, N	oeckite-quartz- o. 9).	syenite, Kogaı	a Zan, analyst	T. Miyashiro-]	LNOUE (MIS	rashrko, A. and
No. 3.	Riebeckite froi MIYASHIRO, T.,	m quartz-syeni 1956, p. 48, N	te-pegmatite, o. 10).	Kogan Zan, aı	nalyst S. Komar	zu (Harada 19	39, ref. MI	vasureo, A. and
No. 4.	Riebeckite fron	n nepheline-sye	nite, Mariupol	AINBERG 193	0, p. 94 or 606, N	Vo. a).		
No. 5.	Riebeckite from	m riebeckite-fe	lsite, North R	oe, analyst C.	O. HARVEY (PHI	EMISTER et al.,	1950, p. 368	, No. 1).
No. 6.	Magnesioriebec	kite from crys	talline schists,	Bizan, analyst	H. HARAMURA	(MIYASHIRO and	I IWASAKI 1	957, p. 701).
No. 7.	Crocidolite, Co	chabamba, ana	lyst H. B. WII	ik (Ernst, 196	0, p. 36, see also	WHITTAKER 194	l9).	
No. 8.	Amphibole fro	m albite-vein, <sup>1</sup>	7als, analysts '	Гн. Hücı and ]	E. SPYCHER, this	paper.		

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<b>00</b>	37 78	10.63	8 56	24.46	0.67	0.51	8.02	2.99	2.06	0.73	n. d.	3.13	0.00	n d		99.54
7	34.41	9.16	5.10	26.07	0.34	2.13	8.98	1.41	1.67	5.94	n. d.	3.52	0.10	n. d.		99.83
9	35.42	8.89	9.73	24.48	1.17	0.17	6.93	5.13	3.23	1.34	n. d.		3.15	n. d.		99.64*)
ĩÇ	36.86	12.10	7.41	23.35	0.77	1.90	10.59	3.20	1.20	n. d.	n. d.	0.60	0.70	n. d.		98.68
4	38.50	10.88	6.70	27.28	tr.	1.40	11.30	1.22	1.66	tr.	n. d.	1.27	0.00	n. d.		100.21
3	37.40	12.34	4.16	25.84	1.24	2.20	9.72	1.80	1.36	3.20	n. d.	0.00	0.60	n. d.		99.86
61	37.49	10.81	7.52	25.14	0.95	1.35	9.77	2.06	1.91	0.86	n. d.	10.6	2.01	n. d.		99.87
I	34.184	11.527	12.621	21.979	0.629	1.353	9.867	3.290	2.286	n. d.	n. d.	010.0	0.940	n. d.		98.084
	SiO <sub>2</sub>	$AI_2O_3$	$Fe_2O_3$	FeO	MnO	MgO	CaO	$Na_2O$	$K_2O$	$\Gamma iO_2$	$P_2O_5$	$H_2O^+$	H20-	Ē <b>r</b> u	F = 0	

Table 4. Chemical analyses and optical properties of some ferrohastingsites

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X-7	0.027	0.017	0.024	0.020	n. d.	0.022	0.021	0.017 - 0.011
X	1.705	1.697	1.698	1.693	n. d.	1.699	1.693	1.695 - 1.682
Τ	1.731	1.713	1.719	1.710	n. d.	1.719	n. d.	n. d.
Z	1.732	1.714	1.722	1.713	n. d.	1.721	1.714	1.711 - 1.693
X	yellow	yellow	yellow	green-	light		yellow-	greenish
				$\mathbf{brown}$	olive gr.		green	yellow
у	deep gr.	blue-gr.	olive	brown-	yellow-		olive	olive brown-
	blue		green	green	green		brown-gr.	green
2	deep ol.	olive-	deep	blue-	blue-		dark	dark
	green	green	green	green	green		green	green
ext.	Y/c 13°	$Y/c 15^{\circ}$	$Z/c 20^{\circ}$	Z/c 17°	$Z/c 9^{\circ}$	$\mathbf{Z}/c$ 20°	Z/c 12°-13°	$Y/c 15^{\circ}-27^{\circ}$
2 V	$25^{\circ}$	$16^{\circ}$	47°	small	n. d.	36°	36°	0°-35°
disp.	$r {\ll} v$	r≪v	r > v	n. d.	n. d.	n. d.	r>v	r < v
¦ = q	$\mathbf{p} = \mathbf{Z}$	$\mathbf{b} = \mathbf{Z}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Z}$
	X > Y > X	Y > Z > X	$\mathbf{Y} > \mathbf{Z} > \mathbf{X}$	Z > Y > X	$Z \ge Y > X$	Y ≧ Z > X	$Z \ge Y > X$	$Z \ge Y > X$
S. G.	n. d.	n. d.	n. d.	n. d.	n. d.	n. d.	n. d.	n. d.
No. 1. F	<sup>T</sup> errohastingsite	e from nepheli	ne-syenite, Dur	ıgannon, ADAM	is and Harrin	GTON 1896 (B	ILLINGS 1928,	p. 288, No. a).
No. 2. 1	Perrohastingsite	e from nephelir	re-syenite, Alm	unge, QUENSEL	, 1914 (BILLING	s 1928, p. 288,	No. b).	
No. 3. 1	<b>Terrohastingsit</b> (	e from nordma	rkite, analyst V	V. H. HERDSMA	N (BILLINGS 19	<b>)28, p. 288, No</b> .	. c).	
No. 4. I	<sup>r</sup> errohastingsite	e from metamo	rphic limestone	», Custer Cy, a	nalyst E. V. Sr	iannon (Billi	NGS 1928, p. 20	88, No. d).
No. 5. I	<b>Hudsonite from</b>	ı pegmatite, Co	rmwall, analyst	S. WEIDMAN I	903 (BILLINGS	1928, p. 288, N	lo. e).	
No. 6. I	Terohastingsite	from alkalisye	nite, Beverly (V	Ивіент 1900).				

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No. 7. Ferrohastingsite from pyroxene-granite, Mariupol (AINBERG 1930, p. 87 or 599, No. a).

No. 8. Ferrohastingsite from mariupolite, Mariupol (AINBERG 1930, p. 89 or 601, No. a).

\*) Sum is 98.64 instead of 99.64 as given in AINBERG's paper.

15	49.22	7.78	8.28	10.49	0.23	10.52	7.45	3.54	0.26	0.23	n. d.	1.80	0.32	n. d.			100.12
14	39.28	11.73	6.77	22.85	0.20	2.78	11.62	0.50	1.69	0.73	n. d.	1.97	0.12	0.52	100.76	0.22	100.54
13	37.33	12.78	8.58	21.65	1.36	1.48	9.46	3.49	1.89	1.32	n. d.	0.81	0.22	n. d.			100.37
12	38.89	10.22	9.53	21.40	0.48	2.23	9.08	3.78	0.97	1.43	n. d.	1.84	n. d.	n. d.			99.85
11	38.03	11.59	6.81	23.72	1.11	2.87	9.75	2.30	1.90	0.22	n. d.	1.20	n. d.	0.05			99.55
10	38.89	11.88	7.56	21.88	0.66	2.68	10.66	1.75	1.99	0.65	n. d.	1.19	(0.02)	0.21			100.00
6	37.51	8.50	11.41	23.21	1.70	0.56	1.91	3.79	2.10	0.73	n. d.	2.32	0.33	0.14	100.21	0.06	100.15
88	38.82	9.07	5.57	25.87	0.25	1.91	11.79	2.79	0.92	0.58	n. d.	2.31	0.27	0.06	0.20	0.07	100.34
	$SiO_2$	$Al_2O_3$	$\rm Fe_2O_3$	FeO	MnO	MgO	CaO	$Na_2O$	$K_2O$	$TiO_2$	$P_2O_5$	$H_2O^+$	$H_{s}O^{-}$	Ē	CI	$\mathbf{F} = \mathbf{O}$	

Table 4 cont.

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0.010-0.016 1.66 -1.67 n. d. 1.675-1.685 c = blue	b = olive green a = yellowish Y/c 20-30° Z/c 12-20°	$0^{\circ}-48^{\circ} (x)$ b = Z  or  Y c > b > a 3.31	(Miyashiro, A.
0.019 1.688 1.707 1.716 pale vellow	$f_{\rm cont}$ brown green br. green- green blue $Z/c 21^{\circ}$ $46^{\circ}$	b = Y Y > Z > X	or 601, No. a.I) H. Yoshizawa
0.018 1.706 1.723 1.724	20°	3.458	(МАТЕТМОТО 1960). TCZ 1925 (АІМЕЕВС 1930, p. 89 SLIE 1945, p. 85, No. IV). 8 (FosllE 1945, p. 85, No. V). 80 (FosllE 1945, p. 85, No. VII). syenite, Fukushinzan, analyst 1 syenite, Fukushinzan, analyst 1
0.018 1.694 1.710 1.712 green- vellow	dark ol. dark ol. dark blue- green $Z/c 171/_2^{\circ}$	r > v $b = Z$ $3.390$	Ubira mine, Kyushu, Japan syenite, Mariupol, Moreozew nite, analyst M. KLÜVER (Fo ite, Österkär, P. GELJER 1913 tie, Österkär, P. GELJER 1913 ains, analyst W. KUNITZ 193 hastingsite-biotie-nepheline- p. 48, No. 7).
light brownish	green dark gree green- brown		sute from the om nepheline. sute from gran from pegmati White Mount from aegurine Bo, T., 1956,
0.018 1.695 1.710 1.713 greenish brown	$dive olive green bluish green Z/c 16\frac{1}{20^{\circ}}$	3.32	Ferrohasting Taramite fre Ferrohasting Hornblende Hastingsite, Hastingsite and MIYASHI
X-Z X X x x x	y z ext. 2 V	disp b = ? S. G.	No. 8a. No. 10. No. 11. No. 12. No. 13.

No. 15. Amphibole from albite-vein, Vals, analysts TH. HÜGI and E. SPYCHER, this paper. A. and MIYASHIRO, T., 1956, p. 48, No. 8).

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	9	49.22	7.78	8.28	10.49	0.23	10.52	7.45	3.54	0.26	0.23	n. d.	1.80	0.32	n. d.	100.12
	ນົ	43.30	10.69	3.94	7.00	0.35	16.02	9.73	4.58	0.66	1.55	n. d.	1.80		<b>n.</b> d.	99.62
	4	39.23	14.38	2.92	8.56	0.65	13.01	11.70	3.05	0.98	4.53	n. d.	0.36		n. d.	99.37
	eo	39.23	11.94	5.92	12.04	0.10	11.01	11.21	2.33	0.99	7.99	n. d.	0.36		n. d.	99.54*)
	2	38.633	11.974	3.903	11.523	0.729	10.200	12.807	3.139	1.489	5.035	n. d.	n. d.	n. d.	n. d.	99.432
	1	38.04	13.50	6.21	15.85	1.21	7.26	12.42	3.21	1.68	1.06	n. d.	n. d.	n. d.	n. d.	100.44
a		$SiO_2$	$Al_2O_3$	$Fe_2O_3$	FeO	MnO	MgO	CaO	$Na_2O$	$K_2O$	$TiO_2$	$P_2O_5$	$H_2O^+$	$H_2O^-$	Γщ	

 Table 5. Chemical analyses and optical properties of some femaghastingsites and magnesiohastingsites

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Z-X	0.021	0.033	0.029	0.029	0.017	0.010 - 0.016
X	1.679	1.670	1.676	1.676	1.653	1.66 - 1.67
Χ	1.694	1.692	1.692	1.692	1.663	n. d.
Z	1.698	1.703	1.705	1.705	1.670	1.675-1.635
×	light brown	light brown	yellow	yellow	pale yellow	c = blue
y.	$\mathbf{brown}$	brown	deep brown	deep brown	green	b = olive green
2 13	green brown	brown	deep brown	brown	green	a = yellowish
		hint of gr.	I			
ext.		Z/c 15°	$Z/c 16^{\circ}$	$Z/c 16^{\circ}$	$\mathbf{Z}/\mathbf{c} = 40^{\circ}$	$Y/c \ 20^{\circ}-30^{\circ}$
	-					Z/c 12°-20°
2 V	$60^{\circ}$	80° .	$76^{\circ}$	$16^{\circ}$	$64^{\circ}$	$0^{\circ}-48^{\circ}$ (x)
disp.	r < v	r < v	$\Gamma > V$	r > v	low	
ہ 1 ' ا	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Y}$	$\mathbf{b} = \mathbf{Y}$	b = Z  or  Y
S. G.	3.518	n. d.	n. d.	3.159	3.160	3.31
	Y > Z > X	Y > Z > X	$\mathbf{Y} = \mathbf{Z} > \mathbf{X}$	$\mathbf{Y} = \mathbf{Z} > \mathbf{X}$	n. d.	c > b > a
No. 1-6.	BILLINGS 1928, p. 29(	0, No. f, g, h/, h <sub>2</sub> , i.				
No. 1. Fe	maghastingsite from	akerite, Cuttingsvil	le.			
No. 2. Fe	maghastingsite from	essexite, Mt. Johns	on.			
No. 3. Tii	taniferous femaghast	ingsite from diorite,	Montreal. NiO is 0	0.04, SrO is 0.01.		
No. 4. M£	agnesiohastingsite fro	m essexite, Montree	al.			
No. 5. Mt	agnesiohastingsite fro	m metamorphosed	limestone, Iron Hill			
Ż.	B. The optimal prope	stries of No. 4 and 5	are not determined	d on the analyzed r	naterial.	
No. 6. AI	nphibole from albite	vein, Vals, analysts	Тн. Hüer and E. S	SPYCHER, this pape	г.	

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\*) Sum is 102.12 instead of 99.54 as given in BILLINGS' paper.

higher figure than was found in the amphibole under discussion. These calculations also show that the analyses given by BILLINGS do not all lead to a satisfactory solution in terms of endmembers. The first analysis might even be wrong, the sum of Al and Si being smaller than 8 and the amount of Ca + Na + K being larger than 3. The same applies to analysis no. 5, which can only be brought to a matching set of endmembers by using SHîDO's variant. The last table, 5, compares the analyses of femaghastingsites and magnesiohastingsites given by BILLINGS with those of our amphibole, showing that here too, disregarding the printer's error in no. 3, a considerable difference exists.

Summarizing, the amphibole under consideration may not be called a riebeckite or a hastingsite. This leaves us with the problem of finding a name for an amphibole that is of widespread occurrence in the metamorphic rocks of certain parts of the Swiss Alps, and is also reported from other regions.

Formerly these amphiboles were called hastingsites on the assumption that similarity of optical properties indicates similarity of chemical composition. This, however, we have shown to be a *fallacy*. In the past the problem of a name would have been solved by proposing a new one for this mineral. The confusion in the nomenclature of amphiboles, is, however, so enormous, that only specialists could find a way out of this chaos. The authors, prefer, therefore, to wait until a sound proposal for a reclassification of amphiboles is made, based on chemical analysis, X-ray data and optical properties. Such a reclassification must necessarily take into account genetical relationships. For the time being the amphibole should be referred to as a *ferrian sodium-amphibole*.

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